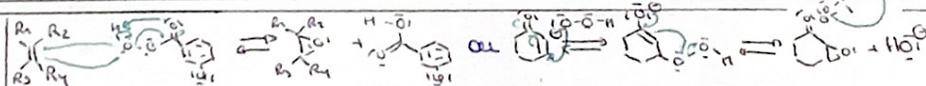
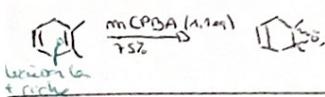


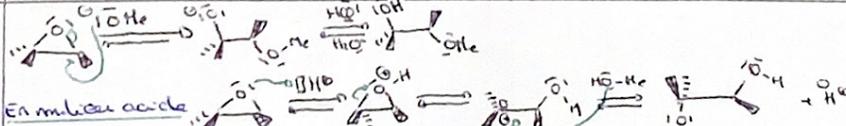
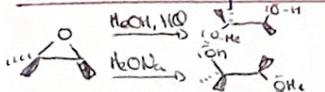
# Oxydation

	Nom	Alcool primaire	Alcool secondaire	Mécanisme
$K_2Cr_2O_7, H_2SO_4$ diluée	/	acide carboxylique (eau conc. & milieu)	cétone	quel que soit le solvant :
$CrO_3, H_2SO_4$ diluée	Jones	acide carboxylique (eau conc. & milieu)	cétone	5. milieu aqueux :
2 pyr. $CrO_3$	Sarett (Collins $CrCl_2$ )	aldéhyde	cétone	
Pyr- $H^+$ , $CrO_3$ (PCC - Pyridinium chloro chromate)	Corey Suggs	aldéhyde	cétone	
$(Pyr-H^+)_2 Cr_2O_7^{2-}$ (PDC : pyridinium dichromate)	Cornforth	aldéhyde	cétone	
Chlorure d'oxalyle $NET_3, DMSO$	Swern	aldéhyde	cétone	voir fiche mécanisme n°1
 DMAP	Dem-Martin	aldéhyde	cétone	
$Pr_4N^+, RuO_4^-$ cat TMA  TMA n-méthyl morpholine-n-oxide	Lay-Griffith	aldéhyde	cétone	

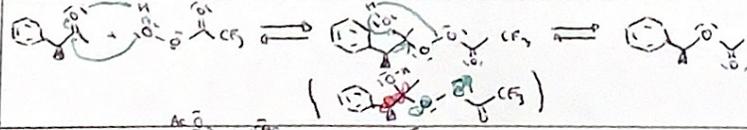
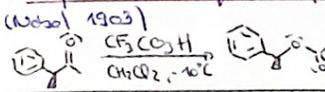
## Époxidation de Prilezhaev



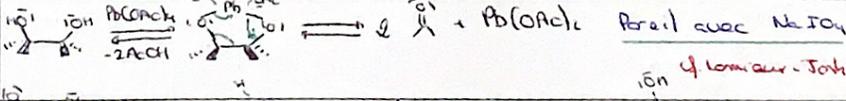
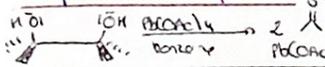
## Ouverture d'époxide



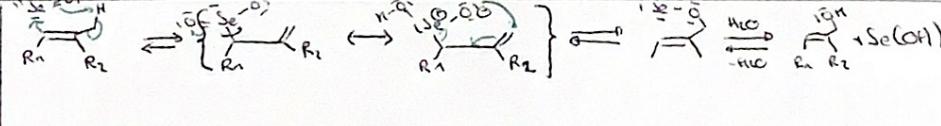
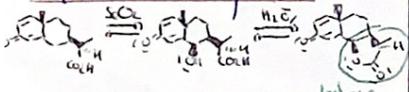
## Oxydation de Baeyer-Villiger



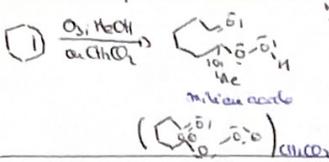
## Oxydation aux plomb



## Oxydation de Riley

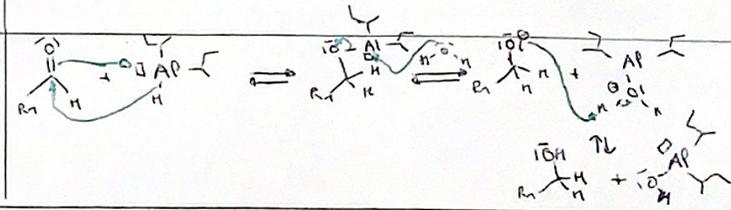
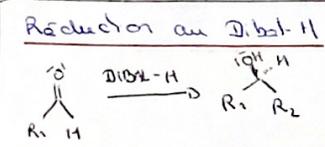
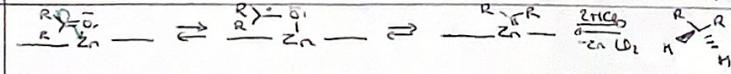
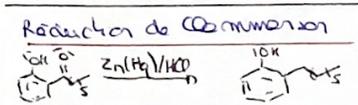
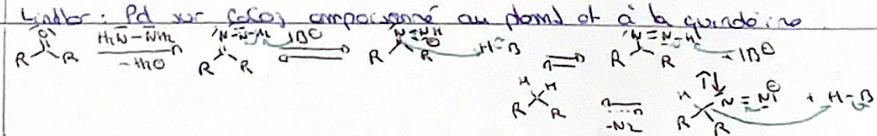
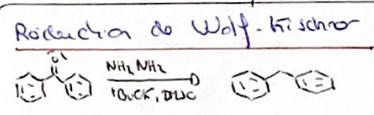
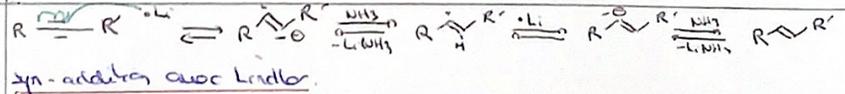
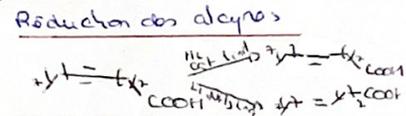
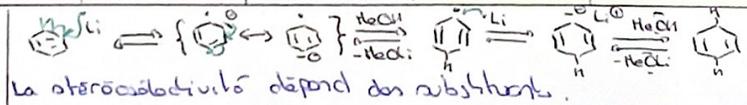
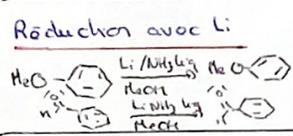


Ozonolyse (1975, Harries and Criegee)



Réduction

Réducteurs	Ester	Aldéhyde	Cétone	Acide	Amide	Nitrile	Imine
LiAlH <sub>4</sub>	alcool	alcool	alcool	alcool (Bnt)	amine (aldéhyde etc)	amine	amine
LiBH <sub>4</sub>	alcool	alcool	alcool	X	X	X	amine
NaBH <sub>4</sub> (s. les esters des acides etc)	alcool (Bnt)	alcool	alcool	X	X	X	amine
Dibal-H	THF/Et <sub>2</sub> O: alcool CH <sub>2</sub> Cl <sub>2</sub> , -78°C: ald.	alcool	alcool	alcool (Bnt)	amine (aldéhyde -78°C)	aldéhyde -78°C	amine
NaBH <sub>3</sub> CN	X	alcool (Bnt)	alcool (Bnt)	X	X	X	amine → sels de f
BH <sub>3</sub>	alcool (Bnt)	alcool (Bnt)	alcool (Bnt)	alcool	amine	amine	amine



Réaction de Staudinger

